pH = 4   First   solution   Second   solution

                          NaOAc   0.1M   HOAc   0.1M

                              ml - ?   ml - ?
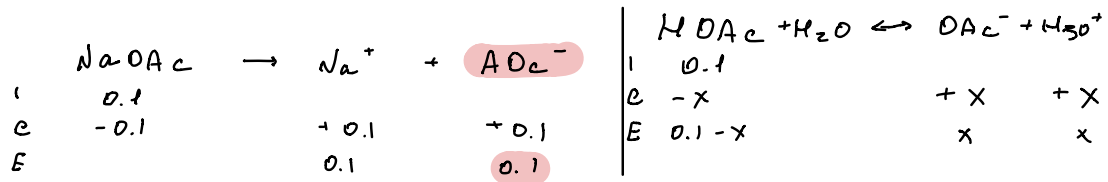
           when   combend   to   one   solution   it   is   20 ml

- log(x) = 4

$1 \cdot 10^{-4}$ — consentration H^+
in the   same   solution (the final   one there   is $1 \cdot 10^{-4} \cdot 0.02$
$2 \cdot 10^{-6}$ mol H^+

                                        | HOAc $+H_2O$ $\longleftrightarrow$ OAc^- $+H_3O^+$
    NaOAc $\longrightarrow$ Na^+ + AOc^- | I   0.1
  I   0.1 | e   -x   +x   +x
  e   -0.1   + 0.1   + 0.1 | E   0.1-x   x   x
  E   0.1   0.1 |

                                              $ka = 1.58 \cdot 10^{-5}$

                                    or - ?

                              HOAc $+H_2O$ $\longleftrightarrow$ OAc^- $+H_3O^+$
                          I   0.1   0.1
                          e   -x   +x   +x
                          E   0.1-x   x   x